

Define the following key words using the glossary:

Acid:

Alkali:

Corrosive:

pH scale:

How does pH change during a neutralisation reaction? Pg104

Give some examples of where neutralisation is useful:

Write a word equation for the reaction between, Pg 107

Magnesium and sulphuric acid:

Zinc oxide with nitric acid:

Copper oxide and hydrochloric acid:

Describe the difference between a concentrated and dilute solution of acid. Pg 100

What are the risks with using acids and alkalis? Pg100

How can we be safe when using them?

Complete the sentences below choosing the correct word:
Acids taste **sour/feel soapy**. Some acidic and alkaline solution are **corrosive/correlated**. A concentrated solution of an acid is **more/less** corrosive than a dilute solution. A concentrated solution has **fewer/more** acid particles per litre than a dilute solution.

Use the information on page 104/5 to fill in the missing words:

A Base cancels out an acid in a _____ reaction.

An alkali is a soluble _____. You can measure pH with an _____ or a pH probe attached to a _____ logger.

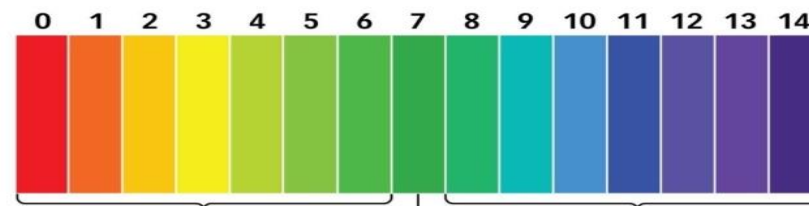
C1. 4 Acids and alkalis

Identify the formulae of three common acids: Pg 106

Describe what is meant by a salt and give some examples:

Describe a method used to produce salt crystals:

Label the pH chart to show where the acids, alkali and neutral is and include some examples of substances. Pg 102



What is an indicator and give some examples of different ones and the results. Pg 102